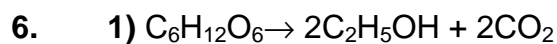


$$\text{x} \Leftrightarrow \text{y} \quad n(\text{y}) = \frac{1}{1} (1 - 0,25) \cdot \frac{1}{1} \cdot 1,25 \text{ mol} \cdot 0,70 = 0,656 \approx \mathbf{0,66 \text{ mol}}$$



2) $m(\text{C}_2\text{H}_5\text{OH}) = 0,750 \text{ cm}^3 \cdot (0,11 - 0,08) \cdot 0,793 \text{ g/cm}^3 = 17,8 \approx \mathbf{20 \text{ g}}$

3) $n(\text{CO}_2) = \frac{2}{2} \cdot 17,8 \text{ g} \cdot \frac{1 \text{ mol}}{46 \text{ g}} = 0,387 \approx \mathbf{0,4 \text{ mol}}$

4) $C(\text{CO}_2, \text{ sampanjas}) = 0,387 \text{ mol} \cdot \frac{0,75}{0,80} \cdot \frac{1}{0,80 \text{ dm}^3} = 0,454 \approx \mathbf{0,5 \text{ mol / dm}^3}$

5) $n(\text{CO}_2, \text{ gaasina}) = 0,387 \cdot \frac{0,05}{0,80} = \mathbf{0,0242 \text{ mol}}$

$$p = \frac{0,0242 \text{ mol} \cdot 0,082 \text{ atm} \cdot \text{dm}^3 \cdot \text{mol}^{-1} \cdot \text{K}^{-1} \cdot 288 \text{ K}}{0,05 \text{ dm}^3} = 11,4 \approx \mathbf{10 \text{ atm}}$$